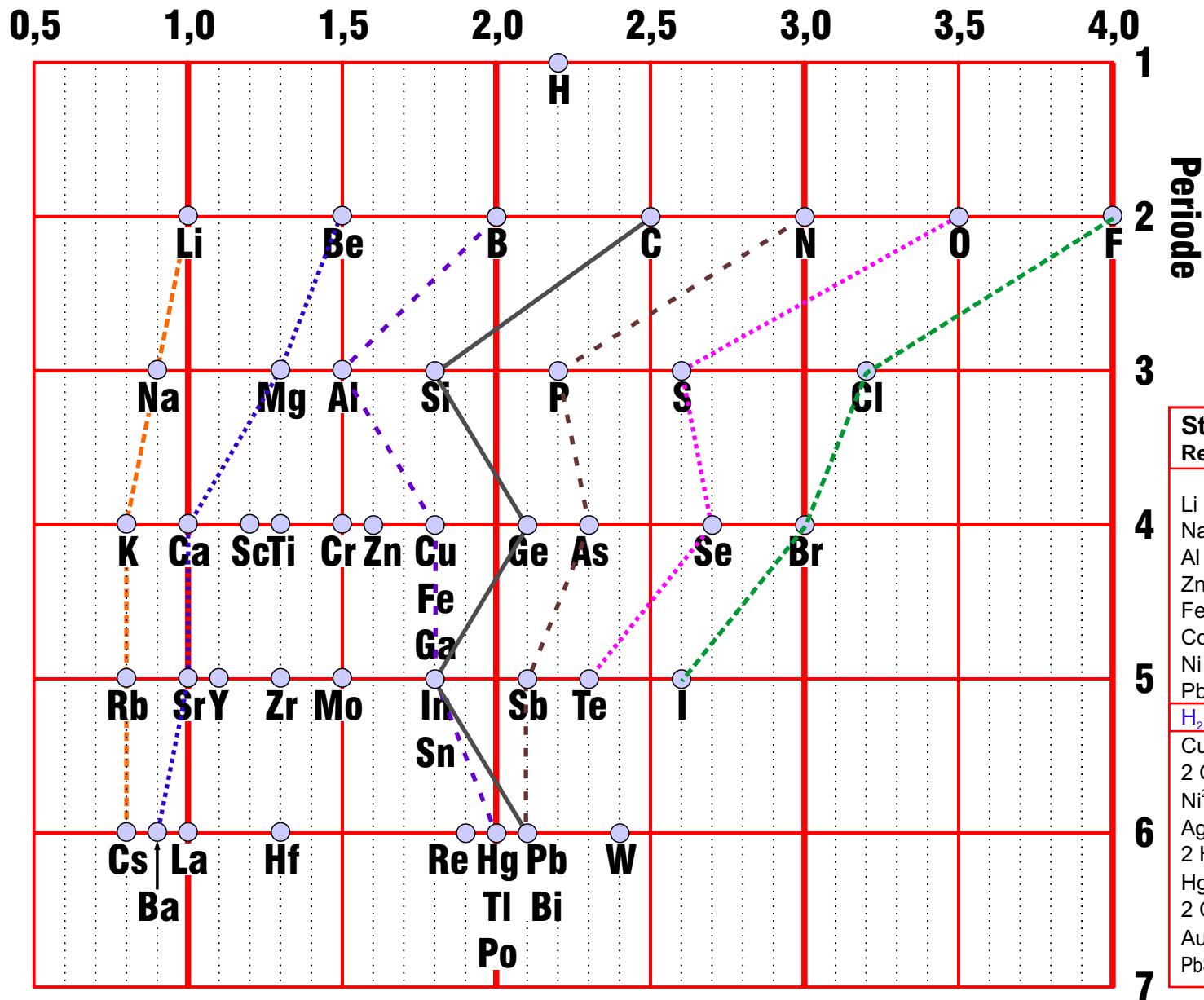


Elektronegativität



Standardpotentiale oder Normalpotentiale		
Reaktionsgleichung:		E° [V]
Li	$\rightleftharpoons \text{Li}^+ + \text{e}^-$	-3,03
Na	$\rightleftharpoons \text{Na}^+ + \text{e}^-$	-2,71
Al	$\rightleftharpoons \text{Al}^{3+} + 3 \text{e}^-$	-1,69
Zn	$\rightleftharpoons \text{Zn}^{2+} + 2 \text{e}^-$	-0,76
Fe	$\rightleftharpoons \text{Fe}^{2+} + 2 \text{e}^-$	-0,44
Cd	$\rightleftharpoons \text{Cd}^{2+} + 2 \text{e}^-$	-0,40
Ni	$\rightleftharpoons \text{Ni}^{2+} + 2 \text{e}^-$	-0,23
Pb	$\rightleftharpoons \text{Pb}^{2+} + 2 \text{e}^-$	-0,13
H_2	$\rightleftharpoons 2 \text{H}^+ + 2 \text{e}^-$	$\pm 0,00$
Cu	$\rightleftharpoons \text{Cu}^{2+} + 2 \text{e}^-$	+0,34
2OH^-	$\rightleftharpoons \frac{1}{2} \text{O}_2 + \text{H}_2\text{O} + 2 \text{e}^-$	+0,40
Ni^{2+}	$\rightleftharpoons \text{Ni}^{3+} + \text{e}^-$	+0,80
Ag	$\rightleftharpoons \text{Ag}^+ + \text{e}^-$	+0,81
$2 \text{H}_2\text{O}$	$\rightleftharpoons \text{O}_2 + 4 \text{H}^+ + 4 \text{e}^-$	+0,82
Hg	$\rightleftharpoons \text{Hg}^{2+} + 2 \text{e}^-$	+0,86
2Cl^-	$\rightleftharpoons \text{Cl}_2 + 2 \text{e}^-$	+1,36
Au	$\rightleftharpoons \text{Au}^{3+} + 3 \text{e}^-$	+1,42
$\text{PbSO}_4 + 2 \text{H}_2\text{O}$	$\rightleftharpoons \text{PbO}_2 + 4\text{H}^+ + \text{SO}_4^{2-} + 2 \text{e}^-$	+1,69